

Name: _____ Date: _____ Block: _____

Worksheet 4-1: Common Acids and Bases

Use Hebden p. 109-114 to fill in the following worksheet for your notes.

1. What is the chemical name for **muriatic** acid? _____
2. Which acid turns skin yellow on contact? _____
3. Which acid is used in car batteries? _____
4. Which acid is found in the human stomach? _____
5. Which acid, in its concentrated form, will corrode copper? _____
6. Concentrated sulphuric acid from the supplier is _____% H_2SO_4 and _____% H_2O . The molar concentration of this acid is _____M.
7. Which acid turns skin white on contact? _____
8. List the four common uses of nitric acid. _____

9. When ammonia (NH_3) gas dissolves in water, it forms a compound called _____

10. What is the chemical name for **caustic potash**? _____
11. What is the chemical name for **caustic soda** or **lye**? _____
12. Which acid is a strong dehydrating agent? _____
13. Concentrated hydrochloric acid from the supplier is _____M or _____% HCl .
14. Give five things that are manufactured with the help of sulphuric acid. _____

15. Which acid has a choking odour? _____
16. What two common bases are used in the production of soaps? _____

17. Give the name of an alkaline gas which is highly soluble in water. _____
18. Nitric acid from the supplier is _____ M or _____ % HNO_3 .
19. Which acid is a **non-electrolyte** in its concentrated form? _____
20. Drain cleaner is made up of mainly _____
21. Name two compounds that absorb CO_2 from the air. _____

22. Which acid is used in the manufacture of textiles? _____
23. _____ acid is used to remove "scale" from boilers.
24. Name an acid which produces a lot of heat when mixed with water. _____
25. Name a base which produces a lot of heat when added to water. _____
26. _____ is used as an electrolyte in alkaline batteries.
27. Write balanced formula equations for the reactions which happen when the following are mixed:
- a) sodium hydroxide and phosphoric acid

- b) sulphuric acid and aluminum hydroxide

- c) $\text{Fe}(\text{OH})_3 + \text{HNO}_3$
